

Please check the examination details below before entering your candidate information

Candidate surname		Other names	
Pearson Edexcel International Advanced Level		Centre Number	Candidate Number

Tuesday 9 October 2018

Morning (Time: 1 hour 30 minutes)

Paper Reference **WCH01/01**
Chemistry
Advanced Subsidiary
Unit 1: The Core Principles of Chemistry
You must have: Scientific calculator

Total Marks

Instructions

- Use **black** ink or **black** ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided
– *there may be more space than you need.*

Information

- The total mark for this paper is 80.
- The marks for **each** question are shown in brackets
– *use this as a guide as to how much time to spend on each question.*
- Questions labelled with an **asterisk** (*) are ones where the quality of your written communication will be assessed
– *you should take particular care with your spelling, punctuation and grammar, as well as the clarity of expression, on these questions.*
- A Periodic Table is printed on the back cover of this paper.

Advice

- Read each question carefully before you start to answer it.
- Show all your working in calculations and include units where appropriate.
- Check your answers if you have time at the end.

Turn over ►

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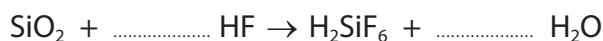
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SECTION A

Answer ALL the questions in this section. You should aim to spend no more than 20 minutes on this section. For each question, select one answer from A to D and put a cross in the box ☐. If you change your mind, put a line through the box ☒ and then mark your new answer with a cross ☐.

- 1 Silicon dioxide reacts with hydrogen fluoride to form water and a compound with the formula H_2SiF_6 .



The mole ratio of HF to H_2O in the balanced equation is

- ☐ A 1:2
☐ B 3:1
☐ C 2:1
☐ D 6:1

(Total for Question 1 = 1 mark)

- 2 For safety reasons, the concentration of lead in paint should not exceed 600 parts per million (ppm) by mass.

Therefore, the mass of lead in one kilogram of paint should not exceed

- ☐ A 0.06 g
☐ B 0.60 g
☐ C 6.0 g
☐ D 60 g

(Total for Question 2 = 1 mark)

- 3 The solution containing the greatest number of chloride ions is

- ☐ A 10 cm^3 of $1.00 \times 10^{-2} \text{ mol dm}^{-3} \text{ AlCl}_3$
☐ B 20 cm^3 of $1.50 \times 10^{-2} \text{ mol dm}^{-3} \text{ MgCl}_2$
☐ C 30 cm^3 of $1.50 \times 10^{-2} \text{ mol dm}^{-3} \text{ HCl}$
☐ D 10 cm^3 of $2.50 \times 10^{-2} \text{ mol dm}^{-3} \text{ CaCl}_2$

(Total for Question 3 = 1 mark)

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4 Which statement is true about the ions $^{55}\text{Mn}^{2+}$ and $^{56}\text{Fe}^{2+}$?

- ☐ A $^{55}\text{Mn}^{2+}$ is deflected less in a mass spectrometer than $^{56}\text{Fe}^{2+}$.
- ☐ B They have the same number of electrons.
- ☐ C $^{55}\text{Mn}^{2+}$ has more protons than $^{56}\text{Fe}^{2+}$.
- ☐ D They have the same number of neutrons.

(Total for Question 4 = 1 mark)

5 10 cm^3 of a $1.00 \times 10^{-2}\text{ mol dm}^{-3}$ solution needs to be diluted to make the concentration $5.00 \times 10^{-4}\text{ mol dm}^{-3}$.

What volume of water, in cm^3 , should be added?

- ☐ A 20
- ☐ B 40
- ☐ C 190
- ☐ D 200

(Total for Question 5 = 1 mark)

6 The Avogadro constant is $6.0 \times 10^{23}\text{ mol}^{-1}$.

The number of **atoms** in 15 g of nitrogen monoxide, NO, is

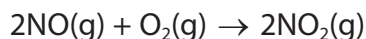
- ☐ A 3.0×10^{23}
- ☐ B 6.0×10^{23}
- ☐ C 2.4×10^{24}
- ☐ D 9.0×10^{24}

(Total for Question 6 = 1 mark)

Use this space for any rough working. Anything you write in this space will gain no credit.



- 7 Nitrogen monoxide reacts with oxygen to form nitrogen dioxide.



200 cm³ of nitrogen monoxide is mixed with 350 cm³ of oxygen.

What is the total volume, in cm³, of the gaseous mixture when the reaction is complete?

All volumes are measured at the same temperature and pressure.

- ☐ A 200
- ☐ B 350
- ☐ C 450
- ☐ D 550

(Total for Question 7 = 1 mark)

- 8 The first six successive ionisation energies of an element X are given in the table.

Ionisation energy	1st	2nd	3rd	4th	5th	6th
Value / kJ mol ⁻¹	789	1577	3232	4356	16 091	19 785

The formula of the oxide of X is most likely to be

- ☐ A XO₂
- ☐ B XO₃
- ☐ C X₂O
- ☐ D X₂O₃

(Total for Question 8 = 1 mark)

- 9 The total number of occupied orbitals in the **third** quantum shell of a silicon atom in its ground state is

- ☐ A 2
- ☐ B 3
- ☐ C 4
- ☐ D 5

(Total for Question 9 = 1 mark)

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10 Which of these statements is correct?

- ☐ A The ionic radii of the alkali metals increase down the group.
- ☐ B The ionic radii for the ions Na^+ , Mg^{2+} , Al^{3+} increase across this series.
- ☐ C The first ionisation energies of the alkali metals increase down the group.
- ☐ D The melting temperatures of successive elements in Period 3 always increase across the period.

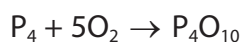
(Total for Question 10 = 1 mark)

11 Which compound would be expected to show the greatest covalent character?

- ☐ A LiBr
- ☐ B LiI
- ☐ C KF
- ☐ D KCl

(Total for Question 11 = 1 mark)

12 Phosphoric(V) acid, H_3PO_4 , can be made from phosphorus in two stages.



Data

Formula	P_4	O_2	P_4O_{10}	H_2O	H_3PO_4
Molar mass / g mol^{-1}	124	32	284	18	98

The percentage atom economy, by mass, for the production of phosphoric(V) acid from phosphorus is

- ☐ A 58.0
- ☐ B 69.0
- ☐ C 72.4
- ☐ D 100

(Total for Question 12 = 1 mark)



13 This question is about the reaction of nickel(II) carbonate and hydrochloric acid.



(a) The ionic equation for this reaction is

(1)

- ☐ **A** $\text{NiCO}_3(\text{s}) + 2\text{H}^+(\text{aq}) \rightarrow \text{Ni}^{2+}(\text{aq}) + \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{l})$
- ☐ **B** $\text{Ni}^{2+}(\text{s}) + 2\text{Cl}^-(\text{aq}) \rightarrow \text{NiCl}_2(\text{aq})$
- ☐ **C** $\text{Ni}^{2+}(\text{s}) + 2\text{HCl}(\text{aq}) \rightarrow \text{NiCl}_2(\text{aq}) + 2\text{H}^+(\text{aq})$
- ☐ **D** $\text{NiCO}_3(\text{s}) + 2\text{HCl}(\text{aq}) \rightarrow \text{Ni}^{2+}(\text{aq}) + 2\text{Cl}^-(\text{aq}) + \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{l})$

(b) Excess hydrochloric acid reacts with 0.20 mol of nickel(II) carbonate.
What is the volume, in dm^3 , of gas produced at room temperature and pressure?
(1 mol of any gas occupies 24 dm^3 at room temperature and pressure)

(1)

- ☐ **A** 1.2
- ☐ **B** 2.4
- ☐ **C** 4.8
- ☐ **D** 9.6

(c) What is the minimum volume of hydrochloric acid with a concentration of 4.0 mol dm^{-3} that reacts with 0.20 mol of nickel carbonate?

(1)

- ☐ **A** 20 cm^3
- ☐ **B** 50 cm^3
- ☐ **C** 100 cm^3
- ☐ **D** 200 cm^3

(Total for Question 13 = 3 marks)

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- 14** When 100 cm^3 of 2.0 mol dm^{-3} sodium hydroxide solution is added to 100 cm^3 of 2.0 mol dm^{-3} sulfuric acid (an excess) to form sodium sulfate, the temperature rise is 12.5°C .

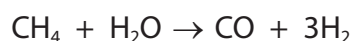
Energy transferred (J) = mass $\times$ 4.2 $\times$ temperature change

What is the enthalpy change of the reaction in kJ mol^{-1} ?

- ☐ **A** $\Delta H = -\frac{200 \times 4.2 \times 12.5}{0.4}$
- ☐ **B** $\Delta H = -100 \times 4.2 \times 12.5 \times 0.4$
- ☐ **C** $\Delta H = -\frac{200 \times 4.2 \times 12.5}{0.2}$
- ☐ **D** $\Delta H = -100 \times 4.2 \times 12.5 \times 0.2$

(Total for Question 14 = 1 mark)

- 15** Hydrogen is manufactured using the reaction



The percentage yield of hydrogen in this process is 90%.

The mass of hydrogen, in tonnes, which can be produced from 160 tonnes of methane is

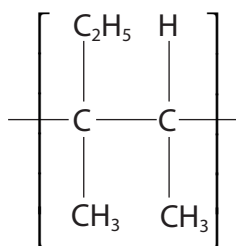
- ☐ **A** 27
- ☐ **B** 54
- ☐ **C** 60
- ☐ **D** 67

(Total for Question 15 = 1 mark)

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16 The repeat unit of a polymer is shown.



What is the systematic name of the monomer which forms this polymer?

- ☐ A 2-ethylbut-2-ene
- ☐ B 2,3-dimethylbut-1-ene
- ☐ C 2-ethylpent-2-ene
- ☐ D 3-methylpent-2-ene

(Total for Question 16 = 1 mark)

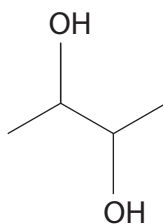
17 A compound contains 31.25% Ca, 18.75% C and 50.00% O.

Its empirical formula is

- ☐ A CaC_2O_4
- ☐ B Ca_2CO_3
- ☐ C Ca_2CO_2
- ☐ D CaCO_3

(Total for Question 17 = 1 mark)

18 Which reagent reacts with but-2-ene to form the compound with the formula shown?



- ☐ A Water
- ☐ B Sodium hydroxide
- ☐ C Hydrogen peroxide
- ☐ D Acidified potassium manganate(VII)

(Total for Question 18 = 1 mark)

TOTAL FOR SECTION A = 20 MARKS

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SECTION B

Answer ALL the questions. Write your answers in the spaces provided.

- 19 A sample of zinc has the relative atomic mass 65.44. The sample contains four isotopes.

The abundance of three of these isotopes is shown.

Relative isotopic mass	64	66	67
Abundance (%)	49.00	27.90	4.50

- (a) (i) Use these data to calculate the relative isotopic mass of the fourth isotope.

Show your working, and give your answer to an appropriate number of significant figures.

(3)

- (ii) State and explain what difference, if any, you would expect between the **chemical** properties of the lightest and heaviest isotopes of zinc.

(1)

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(b) Isotopic masses are determined using a mass spectrometer. The sample under investigation is first converted into gaseous ions.

(i) Ions then pass through slits in a series of electrically charged plates.

Give **two** reasons for this procedure.

(2)

(ii) State how ions of different mass are separated.

(1)

(iii) The ions eventually produce a current in the detector. Data from the detector are used to produce a mass spectrum.

State how the horizontal axis of a mass spectrum is labelled.
Give your answer in **words**, not symbols.

(1)

(c) Complete the electronic configuration of an atom of zinc using s p d notation.

(1)

$1s^2$



- (d) Describe, with the aid of a diagram, the bonding in a sample of zinc. You should state the attractions which hold the particles together in the solid.

(3)

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(Total for Question 19 = 12 marks)



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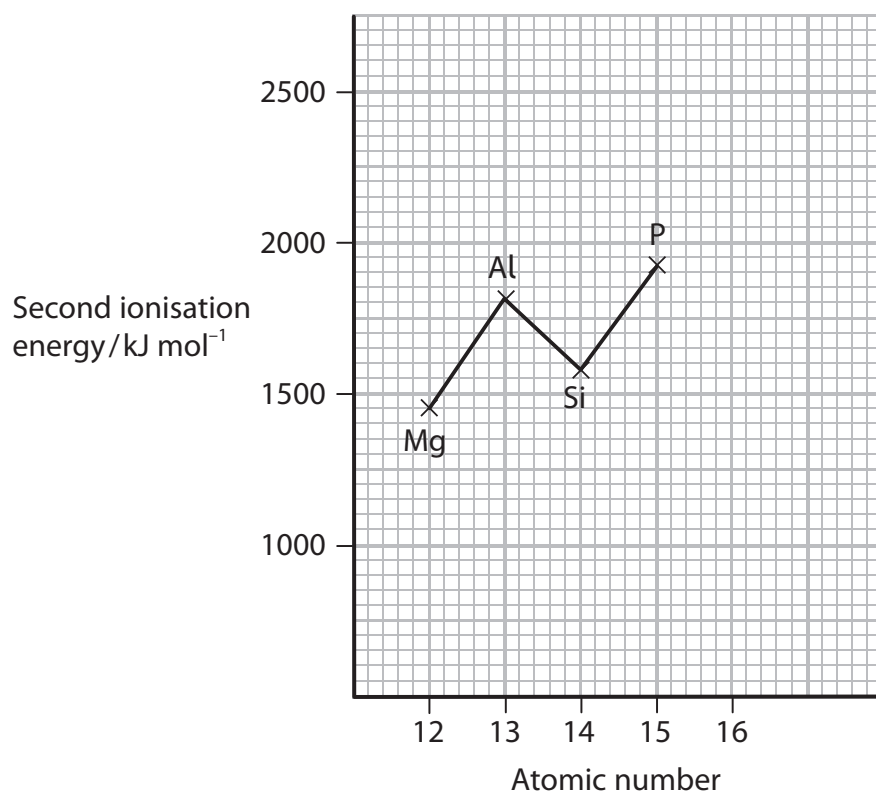
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20 (a) The **second** ionisation energies of some elements in Period 3 are shown on the grid.



(i) Mark on the grid, with a cross, the value you would expect for sulfur.

(1)

(ii) Write an equation, including state symbols, for the **second** ionisation of aluminium.

(2)



- *(iii) Explain why the **second** ionisation energy of aluminium is greater than both the second ionisation energy of magnesium and the second ionisation energy of silicon.

(4)

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- (iv) Predict, with a reason, which element in Period 3 has the highest **second** ionisation energy.

(1)

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(b) Magnesium and sulfur both react with chlorine to form chlorides with a formula XCl_2 .

Magnesium chloride, $MgCl_2$, is ionic. Sulfur dichloride, SCl_2 , consists of covalently bonded molecules.

(i) Describe how the electrical conductivity of these two compounds differs.

(1)

(ii) Draw a dot and cross diagram for sulfur dichloride.

Use crosses (×) for electrons in sulfur and dots (•) for electrons in chlorine.
Only show outer shell electrons.

(2)

(iii) Sketch an electron density map of sulfur dichloride.

(1)

(iv) State how the electron density map of magnesium chloride differs from that of sulfur dichloride.

(1)



(c) The Born-Haber cycle can be used to determine the lattice energy of magnesium chloride.

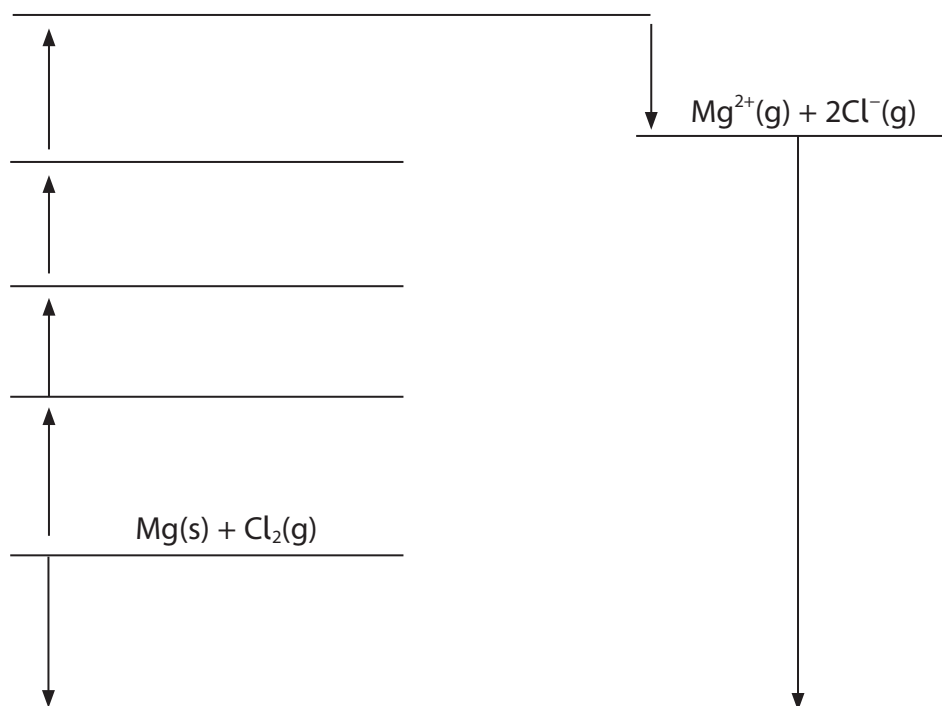
The table below shows the enthalpy changes that are needed.

Energy change	$\Delta H / \text{kJ mol}^{-1}$
Enthalpy change of atomisation of magnesium	+147.7
First ionisation energy of magnesium	+738
Second ionisation energy of magnesium	+1451
Enthalpy change of atomisation of chlorine ($\frac{1}{2}\text{Cl}_2$)	+121.7
First electron affinity of chlorine	-348.8
Enthalpy change of formation of magnesium chloride	-641.3

(i) The diagram shows an incomplete Born-Haber cycle for the formation of magnesium chloride from magnesium and chlorine.

Complete the diagram by writing the **formulae** of the correct species, including state symbols, on the five empty horizontal lines.

(4)



(ii) Calculate the lattice energy of magnesium chloride in kJ mol^{-1} .

(2)

(Total for Question 20 = 19 marks)

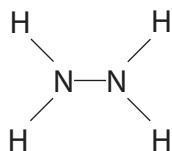
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21 The compound hydrazine, N_2H_4 , is a liquid which is used as a rocket fuel.

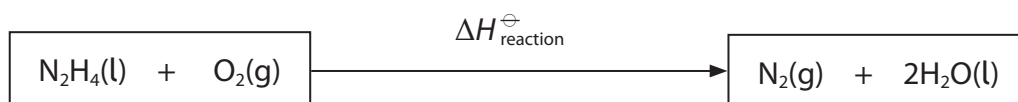


It reacts with oxygen to form nitrogen and water.

- (a) Complete the Hess cycle and, using data in the table, calculate the enthalpy change for the oxidation of hydrazine, $\Delta H_{\text{reaction}}^\ominus$.

Species	Standard enthalpy change of formation / kJ mol^{-1}
$\text{N}_2\text{H}_4(\text{l})$	+50.6
$\text{H}_2\text{O}(\text{l})$	-285.8

(2)



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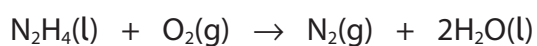
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(b) Some bond enthalpies are given in the table.

Bond	Bond enthalpy / kJ mol^{-1}
N—N	158
O=O	498
N≡N	945
H—O	464
N—H	391

- (i) Calculate the enthalpy change for the oxidation of hydrazine, using the bond enthalpy values in the table.



(3)

- (ii) Give **two** reasons why the enthalpy change calculated using bond enthalpies differs from $\Delta H_{\text{reaction}}^\ominus$ calculated from the Hess cycle.

(2)

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(Total for Question 21 = 7 marks)



22 One component of petrol is decane, $C_{10}H_{22}$.

(a) Decane reacts with chlorine in the presence of ultraviolet light to form a mixture of products.

(i) Complete the equation for the initiation step, including appropriate curly arrows.

(2)



(ii) Write equations, using molecular formulae, for **two** propagation steps.

(2)

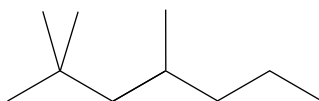
(iii) Write equations, using molecular formulae, for **two** termination steps, other than the one in which chlorine forms.

(2)

(b) The structure of decane can be changed by the process called reforming.

Name the compound shown, which can be produced in this process.

(1)



- (c) Write an equation, using molecular formulae, for the incomplete combustion reaction in which decane reacts to form carbon monoxide and **one** other product.

State symbols are not required.

(1)

- (d) Decane can be cracked to form a mixture of butane, and two different alkenes which have different molecular formulae.

- (i) Write an equation for this reaction, using molecular formulae.
State symbols are not required.

(1)

- *(ii) Explain why geometric isomerism can occur in alkenes and why alkenes produced by this cracking reaction may **not** have geometric isomers.

(2)

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- (iii) Draw the structure of the *trans*, (*E*), isomer of an alkene produced by the cracking reaction in (d)(i).

(1)

(Total for Question 22 = 12 marks)



23 This question is about alkenes.

- *(a) Describe in detail the structure of the $C=C$ double bond in alkenes and hence explain why alkenes are more reactive than alkanes.

(3)

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- (b) Hydrogen bromide reacts with propene to form a mixture of 1-bromopropane and 2-bromopropane.

- (i) Draw the mechanism for the formation of the **major** product in the reaction of propene with hydrogen bromide. You should show relevant dipoles and curly arrows.

(4)

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(ii) State why the amounts of each product are **not** equal.

(1)

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(c) A derivative of propene called allyl bromide, or 3-bromoprop-1-ene, is used to make polymers. The formula of allyl bromide is $\text{CH}_2=\text{CHCH}_2\text{Br}$.

Write the equation for the polymerisation of allyl bromide, showing the structure of the polymer.

(2)

(Total for Question 23 = 10 marks)

TOTAL FOR SECTION B = 60 MARKS

TOTAL FOR PAPER = 80 MARKS



The Periodic Table of Elements

1	2	Key						3	4	5	6	7	0 (8)												
1.0 H hydrogen 1																									
(1)		(2)										(13)		(14)		(15)		(16)		(17)		(18)			
6.9	9.0	relative atomic mass atomic symbol name atomic (proton) number										10.8	12.0	14.0	16.0	19.0	20.2							4.0	
Li lithium 3	Be beryllium 4											B boron 5	C carbon 6	N nitrogen 7	O oxygen 8	F fluorine 9	Ne neon 10							He helium 2	
23.0	24.3											27.0	28.1	31.0	32.1	35.5	39.9							39.9	
Na sodium 11	Mg magnesium 12											Al aluminium 13	Si silicon 14	P phosphorus 15	S sulfur 16	Cl chlorine 17	Ar argon 18							39.9	
39.1	40.1											69.7	72.6	74.9	79.0	79.9	83.8							83.8	
K potassium 19	Ca calcium 20											Ga gallium 31	Ge germanium 32	As arsenic 33	Se selenium 34	Br bromine 35	Kr krypton 36							83.8	
85.5	87.6											114.8	118.7	121.8	127.6	126.9	131.3							131.3	
Rb rubidium 37	Sr strontium 38											In indium 49	Sn tin 50	Sb antimony 51	Te tellurium 52	I iodine 53	Xe xenon 54							131.3	
132.9	137.3											204.4	207.2	209.0	[209]	[210]	[222]							[222]	
Cs caesium 55	Ba barium 56											Tl thallium 81	Pb lead 82	Bi bismuth 83	Po polonium 84	At astatine 85	Rn radon 86							[222]	
[223]	[226]	Elements with atomic numbers 112-116 have been reported but not fully authenticated										[227]	[226]	[225]	[224]	[223]	[222]							[222]	
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												[227]	[226]	[225]	[224]	[223]	[222]							[222]	
												[227]	[226]	[225]	[224]	[223]	[222]							[222]	
												[227]	[226]	[225]	[224]	[223]	[222]							[222]	
												[227]	[226]	[225]	[224]	[223]	[2								